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Poly[bis(μ_4 -benzene-1,4-dicarboxylato)- μ_6 -succinato-dilanthanum(III)]

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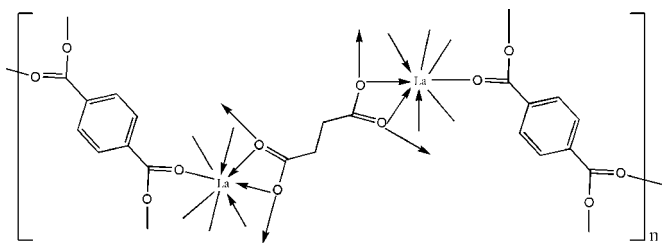
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Key indicators: single-crystal X-ray study; $T = 291$ K; mean $\sigma(\text{C}-\text{C}) = 0.004$ Å; R factor = 0.021; wR factor = 0.051; data-to-parameter ratio = 16.2.

The title compound, $[\text{La}_2(\text{C}_4\text{H}_4\text{O}_4)(\text{C}_8\text{H}_4\text{O}_4)_2]_n$, was synthesized by hydrothermal methods. The La atom is coordinated by four O atoms from four benzene-1,4-dicarboxylate (BDC) ligands and four O atoms from three succinate anions, in a distorted square antiprismatic coordination geometry. The antiprisms are bridged by the benzene-1,4-dicarboxylate and succinate ligands, forming a three-dimensional network. The succinate ion is located on a centre of inversion.

Related literature

For related literature, see: Li & Wang (2005); Li *et al.* (2006); Wang & Li (2005); He *et al.* (2006).



Experimental

Crystal data

$[\text{La}_2(\text{C}_4\text{H}_4\text{O}_4)(\text{C}_8\text{H}_4\text{O}_4)_2]$
 $M_r = 361.06$
 Orthorhombic, $Pbca$
 $a = 14.2014$ (4) Å
 $b = 7.0513$ (1) Å
 $c = 22.1100$ (5) Å

$V = 2214.06$ (9) Å³
 $Z = 8$
 Mo $K\alpha$ radiation
 $\mu = 3.87$ mm⁻¹
 $T = 291$ (2) K
 $0.33 \times 0.27 \times 0.18$ mm

Data collection

Bruker SMART CCD area-detector diffractometer
 Absorption correction: multi-scan (SADABS; Sheldrick, 1996)
 $T_{\min} = 0.293$, $T_{\max} = 0.492$

11927 measured reflections
 2505 independent reflections
 1974 reflections with $I > 2\sigma(I)$
 $R_{\text{int}} = 0.029$

Refinement

$R[F^2 > 2\sigma(F^2)] = 0.021$
 $wR(F^2) = 0.051$
 $S = 1.02$
 2505 reflections

155 parameters
 H-atom parameters constrained
 $\Delta\rho_{\text{max}} = 0.31$ e Å⁻³
 $\Delta\rho_{\text{min}} = -0.53$ e Å⁻³

Table 1

Selected bond lengths (Å).

La—O1	2.4380 (18)	La—O5	2.6244 (19)
La—O2 ⁱ	2.390 (2)	La—O5 ^{iv}	2.532 (2)
La—O3 ⁱⁱⁱ	2.446 (2)	La—O6 ^v	2.573 (2)
La—O4 ⁱⁱⁱ	2.450 (2)	La—O6	2.6884 (19)

Symmetry codes: (i) $-x + 2, -y + 2, -z + 1$; (ii) $-x + \frac{3}{2}, -y + 2, z + \frac{1}{2}$; (iii) $x, -y + \frac{3}{2}, z + \frac{1}{2}$; (iv) $-x + \frac{3}{2}, y + \frac{1}{2}, z$; (v) $-x + \frac{3}{2}, y - \frac{1}{2}, z$.

Data collection: SMART (Bruker, 1998); cell refinement: SMART; data reduction: SAINT (Bruker, 1998); program(s) used to solve structure: SHELXS97 (Sheldrick, 1997); program(s) used to refine structure: SHELXL97 (Sheldrick, 1997); molecular graphics: SHELXTL (Bruker, 1998); software used to prepare material for publication: SHELXTL

Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: BT2565).

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supplementary materials

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Poly[bis(μ_4 -benzene-1,4-dicarboxylato)- μ_6 -succinato-dilanthanum(III)]

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Comment

The title compound, (I), is isostructural with its $[M_2(C_8H_4O_4)_2(C_4H_4O_4)]_n$ [$M = \text{Gd}$ (Wang & Li, 2005), Dy ((Li & Wang, 2005), Nd (Li *et al.*, 2006), Er (He *et al.*, 2006)] analogues. As depicted in Fig.1, The La^{3+} ion is located at the center of a distorted square antiprism geometry and is coordinated by four oxygen atoms from four BDC and four from three succinate anions. The La—O bond distances ranging from 2.390 (2) to 2.6884 (19) Å.

In (I), the succinate ligand is located on an inversion centre and functions as an octadentate ligand, bis-chelating two La atoms with each O atom bridging to another La atom. In this mode, the La atoms are linked into a two-dimensional polymeric sheet parallel to the (001) plane. These sheets are in turn bridged *via* benzene-1,4-dicarboxylate ligands, forming a three-dimensional framework.

Experimental

A mixture of $\text{LaCl}_3 \cdot 6\text{H}_2\text{O}$ (2.00 mmol, 0.70 g), benzene-1,4-dicarboxylic acid (1.0 mmol, 0.16 g), succinic acid (1.0 mmol, 0.10 g), NaOH (6.0 ml, 1 mol/L) and H_2O (20.0 ml) was heated in a 35 ml stainless steel reactor with a Teflon liner at 453 K for 48 h. The column-like crystals were filtered and washed with ethanol. Yield: 14% based on Nd.

Refinement

H atoms were included at calculated positions and treated as riding atoms, with C—H distances of 0.93–0.97 Å and $U_{\text{iso}}(\text{H}) = 1.2U_{\text{eq}}(\text{C})$.

Figures

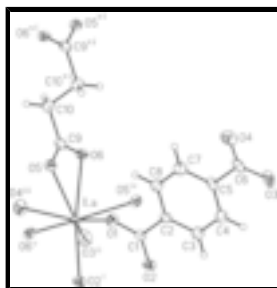


Fig. 1. The coordination environment of the La atom, with the atom-numbering scheme, showing displacement ellipsoids drawn at the 50% probability level. Symmetry codes: (i) $2 - x, 2 - y, 1 - z$; (ii) $3/2 - x, 2 - y, z + 1/2$; (iii) $x, 3/2 - y, z + 1/2$; (iv) $3/2 - x, y + 1/2, z$; (v) $3/2 - x, y - 1/2, z$; (vi) $1 - x, 2 - y, 1 - z$.

Poly[bis(μ_4 -benzene-1,4-dicarboxylato)- μ_6 -succinato-dilanthanum(III)]

Crystal data

$[\text{La}_2(\text{C}_4\text{H}_4\text{O}_4)(\text{C}_8\text{H}_4\text{O}_4)_2]$	$F_{000} = 1368$
$M_r = 361.06$	$D_x = 2.166 \text{ Mg m}^{-3}$
Orthorhombic, <i>Pbca</i>	Mo $K\alpha$ radiation
Hall symbol: -P 2ac 2ab	$\lambda = 0.71073 \text{ \AA}$
$a = 14.2014 (4) \text{ \AA}$	Cell parameters from 168 reflections
$b = 7.0513 (1) \text{ \AA}$	$\theta = 3.3\text{--}26.7^\circ$
$c = 22.1100 (5) \text{ \AA}$	$\mu = 3.87 \text{ mm}^{-1}$
$V = 2214.06 (9) \text{ \AA}^3$	$T = 291 (2) \text{ K}$
$Z = 8$	Column, colorless
	$0.33 \times 0.27 \times 0.18 \text{ mm}$

Data collection

Bruker SMART CCD area-detector diffractometer	2505 independent reflections
Radiation source: fine-focus sealed tube	1974 reflections with $I > 2\sigma(I)$
Monochromator: graphite	$R_{\text{int}} = 0.029$
$T = 291(2) \text{ K}$	$\theta_{\text{max}} = 27.5^\circ$
φ and ω scans	$\theta_{\text{min}} = 1.8^\circ$
Absorption correction: multi-scan (SADABS; Sheldrick, 1996)	$h = -18 \rightarrow 18$
$T_{\text{min}} = 0.293$, $T_{\text{max}} = 0.492$	$k = -9 \rightarrow 5$
11927 measured reflections	$l = -28 \rightarrow 29$

Refinement

Refinement on F^2	Hydrogen site location: inferred from neighbouring sites
Least-squares matrix: full	H-atom parameters constrained
$R[F^2 > 2\sigma(F^2)] = 0.021$	$w = 1/[\sigma^2(F_o^2) + (0.0243P)^2 + 1.3924P]$
$wR(F^2) = 0.051$	where $P = (F_o^2 + 2F_c^2)/3$
$S = 1.02$	$(\Delta/\sigma)_{\text{max}} = 0.030$
2505 reflections	$\Delta\rho_{\text{max}} = 0.31 \text{ e \AA}^{-3}$
155 parameters	$\Delta\rho_{\text{min}} = -0.53 \text{ e \AA}^{-3}$
Primary atom site location: structure-invariant direct methods	Extinction correction: SHELXTL (Bruker, 1998), $F_c^* = kF_c[1 + 0.001x F_c^2 \lambda^3 / \sin(2\theta)]^{-1/4}$
Secondary atom site location: difference Fourier map	Extinction coefficient: 0.00116 (9)

Special details

Geometry. All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

Refinement. Refinement of F^2 against ALL reflections. The weighted R -factor wR and goodness of fit S are based on F^2 , conventional R -factors R are based on F , with F set to zero for negative F^2 . The threshold expression of $F^2 > \sigma(F^2)$ is used only for calculating R -factors(gt) *etc.* and is not relevant to the choice of reflections for refinement. R -factors based on F^2 are statistically about twice as large as those based on F , and R -factors based on ALL data will be even larger.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (\AA^2)

	<i>x</i>	<i>y</i>	<i>z</i>	$U_{\text{iso}}^*/U_{\text{eq}}$
La	0.833810 (10)	1.02208 (2)	0.554317 (6)	0.01592 (7)
O1	0.86146 (16)	1.0440 (3)	0.44571 (8)	0.0277 (5)
O2	1.00219 (14)	0.9173 (3)	0.43107 (8)	0.0279 (5)
C1	0.92144 (19)	0.9655 (4)	0.41250 (12)	0.0198 (6)
C2	0.89650 (19)	0.9240 (4)	0.34781 (12)	0.0225 (6)
C3	0.8095 (2)	0.9761 (5)	0.32544 (14)	0.0326 (8)
H3A	0.7667	1.0383	0.3503	0.039*
C4	0.7857 (2)	0.9356 (5)	0.26579 (14)	0.0359 (8)
H4A	0.7269	0.9697	0.2509	0.043*
C5	0.8497 (2)	0.8446 (5)	0.22869 (13)	0.0289 (7)
C6	0.8238 (2)	0.7940 (5)	0.16472 (14)	0.0341 (8)
C7	0.9370 (2)	0.7938 (5)	0.25116 (13)	0.0361 (8)
H7A	0.9803	0.7333	0.2262	0.043*
C8	0.9603 (2)	0.8325 (5)	0.31042 (13)	0.0337 (8)
H8A	1.0189	0.7971	0.3253	0.040*
O3	0.74602 (16)	0.8515 (4)	0.14369 (9)	0.0424 (6)
O4	0.88149 (19)	0.6945 (4)	0.13534 (10)	0.0486 (7)
O5	0.66405 (13)	0.8794 (3)	0.54300 (10)	0.0285 (5)
O6	0.67149 (13)	1.1776 (3)	0.51616 (9)	0.0247 (4)
C9	0.6242 (2)	1.0327 (5)	0.52874 (16)	0.0293 (7)
C10	0.5157 (2)	1.0426 (5)	0.52970 (16)	0.0376 (8)
H10A	0.4907	0.9706	0.5635	0.045*
H10B	0.4946	1.1729	0.5331	0.045*

Atomic displacement parameters (\AA^2)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
La	0.01657 (9)	0.01669 (11)	0.01450 (9)	−0.00152 (6)	−0.00020 (6)	0.00152 (6)
O1	0.0323 (11)	0.0345 (15)	0.0162 (10)	0.0050 (9)	0.0015 (8)	−0.0020 (9)
O2	0.0243 (10)	0.0378 (14)	0.0217 (10)	−0.0031 (10)	−0.0081 (8)	−0.0016 (9)
C1	0.0228 (14)	0.0197 (17)	0.0169 (13)	−0.0048 (11)	−0.0036 (11)	−0.0002 (11)
C2	0.0240 (14)	0.0248 (18)	0.0186 (14)	−0.0005 (12)	−0.0043 (11)	−0.0025 (12)

supplementary materials

C3	0.0340 (16)	0.039 (2)	0.0246 (16)	0.0136 (14)	-0.0038 (13)	-0.0070 (14)
C4	0.0333 (18)	0.042 (2)	0.0320 (17)	0.0112 (15)	-0.0141 (14)	-0.0037 (16)
C5	0.0432 (19)	0.0241 (19)	0.0195 (14)	-0.0012 (13)	-0.0091 (13)	-0.0032 (13)
C6	0.055 (2)	0.0244 (19)	0.0225 (15)	-0.0097 (15)	-0.0123 (15)	-0.0019 (14)
C7	0.0396 (18)	0.044 (2)	0.0243 (15)	0.0122 (15)	-0.0044 (14)	-0.0143 (15)
C8	0.0294 (16)	0.048 (2)	0.0234 (15)	0.0107 (15)	-0.0088 (13)	-0.0119 (15)
O3	0.0468 (15)	0.0549 (19)	0.0256 (11)	-0.0100 (12)	-0.0167 (10)	0.0095 (11)
O4	0.0760 (18)	0.0420 (18)	0.0279 (12)	0.0110 (14)	-0.0163 (12)	-0.0156 (12)
O5	0.0216 (10)	0.0155 (13)	0.0483 (13)	0.0007 (8)	-0.0087 (9)	0.0021 (10)
O6	0.0236 (10)	0.0165 (12)	0.0339 (11)	-0.0013 (8)	-0.0039 (8)	0.0024 (9)
C9	0.0193 (15)	0.022 (2)	0.0463 (19)	-0.0011 (12)	-0.0118 (14)	0.0020 (15)
C10	0.042 (2)	0.026 (2)	0.045 (2)	0.0062 (15)	-0.0048 (16)	-0.0042 (17)

Geometric parameters (\AA , $^\circ$)

La—O1	2.4380 (18)	C4—C5	1.382 (4)
La—O2 ⁱ	2.390 (2)	C4—H4A	0.9300
La—O3 ⁱⁱ	2.446 (2)	C5—C7	1.383 (4)
La—O4 ⁱⁱⁱ	2.450 (2)	C5—C6	1.504 (4)
La—O5	2.6244 (19)	C6—O4	1.259 (4)
La—O5 ^{iv}	2.532 (2)	C6—O3	1.265 (4)
La—O6 ^v	2.573 (2)	C7—C8	1.378 (4)
La—O6	2.6884 (19)	C7—H7A	0.9300
O1—C1	1.253 (3)	C8—H8A	0.9300
O2—C1	1.265 (3)	O5—C9	1.260 (4)
C1—C2	1.502 (3)	O6—C9	1.254 (4)
C2—C3	1.380 (4)	C9—C10	1.542 (4)
C2—C8	1.386 (4)	C10—C10 ^{vi}	1.512 (7)
C3—C4	1.391 (4)	C10—H10A	0.9700
C3—H3A	0.9300	C10—H10B	0.9700
O2 ⁱ —La—O1	87.98 (7)	C3—C2—C8	119.7 (3)
O2 ⁱ —La—O3 ⁱⁱ	106.10 (7)	C3—C2—C1	120.0 (2)
O1—La—O3 ⁱⁱ	147.90 (8)	C8—C2—C1	120.3 (2)
O2 ⁱ —La—O4 ⁱⁱⁱ	75.13 (8)	C2—C3—C4	120.1 (3)
O1—La—O4 ⁱⁱⁱ	135.66 (8)	C2—C3—H3A	119.9
O3 ⁱⁱ —La—O4 ⁱⁱⁱ	76.39 (9)	C4—C3—H3A	119.9
O2 ⁱ —La—O5 ^{iv}	79.84 (7)	C5—C4—C3	119.9 (3)
O1—La—O5 ^{iv}	80.66 (7)	C5—C4—H4A	120.1
O3 ⁱⁱ —La—O5 ^{iv}	73.90 (8)	C3—C4—H4A	120.1
O4 ⁱⁱⁱ —La—O5 ^{iv}	133.58 (8)	C7—C5—C4	119.8 (3)
O2 ⁱ —La—O6 ^v	103.99 (7)	C7—C5—C6	119.7 (3)
O1—La—O6 ^v	75.02 (7)	C4—C5—C6	120.5 (3)
O3 ⁱⁱ —La—O6 ^v	126.55 (7)	O4—C6—O3	123.8 (3)
O4 ⁱⁱⁱ —La—O6 ^v	70.07 (7)	O4—C6—C5	117.3 (3)
O5 ^{iv} —La—O6 ^v	155.16 (7)	O3—C6—C5	118.9 (3)

O2 ⁱ —La—O5	167.61 (7)	C8—C7—C5	120.3 (3)
O1—La—O5	94.49 (7)	C8—C7—H7A	119.8
O3 ⁱⁱ —La—O5	77.94 (7)	C5—C7—H7A	119.8
O4 ⁱⁱⁱ —La—O5	94.85 (8)	C7—C8—C2	120.1 (3)
O5 ^{iv} —La—O5	112.53 (6)	C7—C8—H8A	119.9
O6 ^v —La—O5	65.15 (6)	C2—C8—H8A	119.9
O2 ⁱ —La—O6	143.61 (7)	C6—O3—La ^{vii}	144.9 (2)
O1—La—O6	78.65 (7)	C6—O4—La ^{viii}	123.0 (2)
O3 ⁱⁱ —La—O6	72.99 (7)	C9—O5—La ^v	152.08 (19)
O4 ⁱⁱⁱ —La—O6	136.12 (8)	C9—O5—La	96.17 (18)
O5 ^{iv} —La—O6	64.75 (6)	La ^v —O5—La	111.17 (7)
O6 ^v —La—O6	104.90 (6)	C9—O6—La ^{iv}	132.9 (2)
O5—La—O6	48.61 (7)	C9—O6—La	93.29 (17)
O2 ⁱ —La—C9	167.94 (8)	La ^{iv} —O6—La	107.90 (7)
O1—La—C9	88.44 (9)	O6—C9—O5	120.9 (3)
O3 ⁱⁱ —La—C9	71.73 (9)	O6—C9—C10	120.1 (3)
O4 ⁱⁱⁱ —La—C9	115.04 (9)	O5—C9—C10	118.9 (3)
O5 ^{iv} —La—C9	88.21 (7)	O6—C9—La	62.32 (15)
O6 ^v —La—C9	86.18 (7)	O5—C9—La	59.41 (15)
O5—La—C9	24.42 (8)	C10—C9—La	168.4 (2)
O6—La—C9	24.40 (7)	C10 ^{vi} —C10—C9	105.4 (3)
C1—O1—La	131.18 (18)	C10 ^{vi} —C10—H10A	110.7
C1—O2—La ⁱ	151.6 (2)	C9—C10—H10A	110.7
O1—C1—O2	123.0 (2)	C10 ^{vi} —C10—H10B	110.7
O1—C1—C2	118.9 (2)	C9—C10—H10B	110.7
O2—C1—C2	118.1 (2)	H10A—C10—H10B	108.8

Symmetry codes: (i) $-x+2, -y+2, -z+1$; (ii) $-x+3/2, -y+2, z+1/2$; (iii) $x, -y+3/2, z+1/2$; (iv) $-x+3/2, y+1/2, z$; (v) $-x+3/2, y-1/2, z$; (vi) $-x+1, -y+2, -z+1$; (vii) $-x+3/2, -y+2, z-1/2$; (viii) $x, -y+3/2, z-1/2$.

Fig. 1

